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## **Review And Reinforcement Chemical Equilibrium**

No. Equilibrium implies

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that the forward and reverse reactions occur at the same rate. 15. The forward and reverse reactions are both occurring, but with no net change in concentration. -16—2 Review Reinforcement, pages 15-16 (H2J[12] - [NH3]IHC1) 9.

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In behavioral psychology, reinforcement is a

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consequence that will strengthen an organism's future behavior whenever that behavior is preceded by a specific ... Supply and demand - Wikipedia, the free encyclopedia

## **16 1 review reinforcement the concept of equilibrium ...**

The basic criterion for equilibrium with a single reaction is:

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$D.G = \sum_i V_i \mu_i$  where  
 $D.G$  is the Gibbs  
function,  $NCOMP$  is the  
number of components  
in the system,  $V_i$  is the  
stoichiometric  
coefficient of species  $i$ ,  
and  $\mu_i$  is the chemical  
potential of species  $i$ .

## **Review of Chemical Equilibria - CaltechAUTHORS**

Atoms tend to share  
electrons in chemical  
bonds. Atoms have 5

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valence electrons.

Chemical  
Never found in nature  
uncombined.

Equilibrium  
Aluminum is in this  
group. Do not normally  
react with other  
elements. Atoms have  
4 valence electrons.

Answer Key  
Most abundant  
element in the Earth's  
atmosphere is in this  
group. Atoms have 6  
valence electrons.

## **REVIEW and REINFORCEMENT Section 6-3**



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Review of Pressure and  
Equilibrium

## Chemical Equilibrium Review - Le Chatelier's and Pressure

Goal: To describe how equilibrium systems can be disrupted and show you how to predict whether certain changes on a system at equilibrium will lead to more products, more reactants, or neither.

Although the concept

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of chemical equilibrium is very important, many reversible reactions in nature never form equilibrium systems.

## **Chapter 16 - The Process of Chemical Reactions**

In a chemical equilibrium, the forward and reverse reactions occur at equal rates, and the concentrations of products and reactants

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remain constant. If we run a reaction in a closed system so that the products cannot escape, we often find the reaction does not give a 100% yield of products.

## **13.1 Chemical Equilibria - Chemistry**

Worksheet 16 -

Equilibrium Chemical  
equilibrium is the state  
where the

concentrations of all

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Equilibrium  
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reactants and products remain constant with time. Consider the following reaction:  $\text{H}_2\text{O} + \text{CO} \rightleftharpoons \text{H}_2 + \text{CO}_2$   
Suppose you were to start the reaction with some amount of each reactant (and no H

## **Worksheet 16 - Equilibrium Chemical equilibrium**

Terms in this set (...)  
chemical equilibrium.  
the state where the

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concentration of all reactants and products remain constant with time. law of mass action.

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equilibrium concentration of each species and component. Mass balances are set to equal the total concentration of each

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component in the system (generally measured analytically), and mass action equations are defined by equilibrium constants ( $K$  or values, also determined experimentally). Thus the mass balances are defined as  $M_j = \sum_i \nu_{ij} x_i$  (4:1)

## **Modelling Chemical Speciation: Thermodynamics, Kinetics**

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LeChatelier's Principle

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16-3 Review and  
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## Equilibrium Answer Key - **AE Equilibrium Answer Keys - Baumritter - Google Sites**

If the chemical equation for a reversible reaction is modified, the equilibrium constant expression must also be modified. If the equation is reversed, the or expression is inverted. If the



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coefficients of the equation are multiplied by a common factor, the or expression is raised to the corresponding power.

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(continued) Solve each of the following problems as directed. Show all your work. 13. How many moles of magnesium are required to react with 2.0 mol of hydrochloric acid?

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Comparing Q vs K  
example. Google  
Classroom Facebook  
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that affect chemical  
equilibrium. Le

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Chatelier's principle. Le  
Chatelier's principle:  
Worked example.  
Introduction to reaction  
quotient  $Q_c$ . The  
reaction quotient  $Q$ .  
Comparing  $Q$  vs  $K$   
example. This is the  
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Key ... Reinforcement,

pages 8-9 . 1. energy

2. radiant 3. chemical

potential 4. joule 5.

created, destroyed ...

17. distillation 18. A ...

equilibrium is

established. For the

example of a liquid

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Chemical

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Chemistry 12 Review  
Sheet on Unit 2

Chemical Equilibrium 1.

What two things are  
equal at equilibrium?

\_\_\_\_\_ and \_\_\_\_\_ 2.

Consider the following  
potential energy

diagram: a) Which



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reaction, forward or reverse, will be affected more by an increase in temperature?

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ecf8427e.