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CHAPTER 13, States of Matter
(continued) SECTION 13.3 THE NATURE OF SOLIDS (pages 396–399) This section describes the highly organized structures of solids, distinguishes between a crystal lattice and a unit cell, and explains how allotropes of an element differ. A Model for Solids (page 396) 1. Is the following sentence true or false?

Name Date Class STATES OF MATTER 13

- b. The minimum temperature at which the substance exists in the liquid state.
- c. The maximum pressure and temperature at which the substance is in equilibrium between the solid, liquid, and gas.
- d. The pressure and temperature at which the substance is in equilibrium between the solid, liquid, and gas states.

Chapter 13 : States of Matter

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13.3 Pascals principle: pressure applied to a confined fluid will be spread throughout the fluid. So $F/A = F'/A'$. A small force applied over a small area can give a bigger force over a bigger area like in a hydraulic system. The pressure you feel under water is due to the weight of the water over you.

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of matter. Whether your reading light comes from a fluorescent lamp or the sun, the light's source is plasma. Although plasma is still somewhat

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mysterious to humans, the plasma state of matter is the most common state in the universe. In this chapter, your exploration of the states of matter will go far beyond everyday, casual observations.

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all matter consists of tiny particles that are in constant motion: normal boiling point: the boiling point of a liquid at a pressure of 101.3 kPa: phase diagram: graph representing the relationships among the solid, liquid, and vapor states of a substance in a sealed container: standard atmosphere: pressure required to support 760 mm of mercury at 25°C

Quia - Chapter 13 "States of Matter"

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Section 13.1 The Gas Laws 1. pressure
2. volume 3. temperature 4.
temperature 5. volume 6. pressure 7.
pressure 8. temperature 9. volume 10.
increases 11. stays the same 12.
decreases 13. increases 14. decreases
15. increases Section 13.2 The
Combined Gas Law and Avogadro's
Principle 1. Gay-Lussac's 2. Boyle's 3.
Charles's

Study Guide - Chapter 12 - States of Matter

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Chemistry Chapter 13: States Of Matter Review . Chemistry Chapter 13: States Of Matter Review Your answer format should be _____, _____ (First word comma second word) 9. The rate of diffusion mainly depends on the mass of the particles involved. A. True. B. False. 10. Smaller particles diffuse slower than heavier particles. A. True. B ...

Chemistry Chapter 13: States Of Matter Review - ProProfs Quiz

28. Use the chart to identify the state of matter described by the following. Many of these have more than one answer! (Use S, L or G in the spaces.) L not easily

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compressible rigid — particles locked into place flows easily Compressible tots of free space between particles does not flow easily

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13.4 Changes of State 13 - Henry County School District

CHAPTER 13 PRACTICE TEST Matching: Match the lettered elements with the numbered elements. Answers may be used more than once. Only one answer may be used for each question. Place all answers on answer sheet provided. (2 point each = 30)

a. decrease(s) f. plasma k. vapor pressure
b. increase(s) g. evaporation l. thermometer
c. no h. phase change m.

Chapter 13 Practice Test 2014 + Answers - CHAPTER 13 ...

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Chapter 13 - States of Matter - 13.3 The Nature of Solids - 13.3 Lesson Check - Page 434: 24 Answer Solids have a definite shape and a definite volume because of the arrangement of the particles.

Chemistry (12th Edition) Chapter 13 - States of Matter ...

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read about gas pressure. Circle the letter of the choice that best completes the statement or answers the question. 13. Pressure is defined as force per unit

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TEACHER GUIDE AND ANSWERS

Chemistry: Matter and Change Teacher Guide and Answers 8 8. c 9. a 10.

temperature and pressure 11. a. vapor

b. solid c. liquid 12. the normal freezing point of water 13. 100.00 degrees

Celsius 14. the temperatures and pressures at which solid water and water vapor could exist

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